

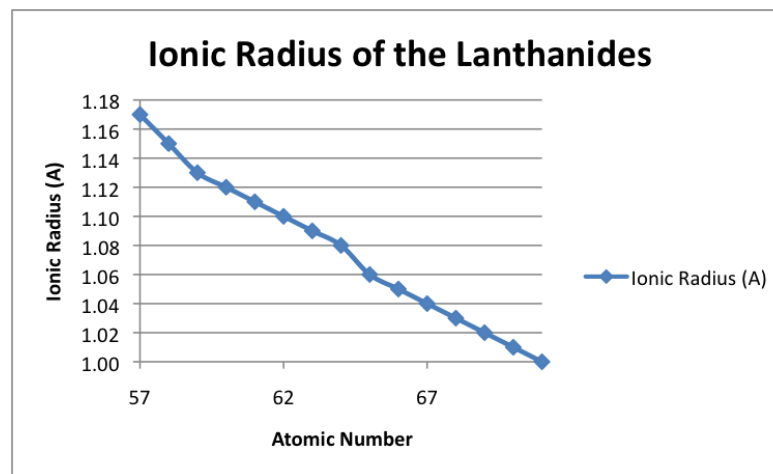
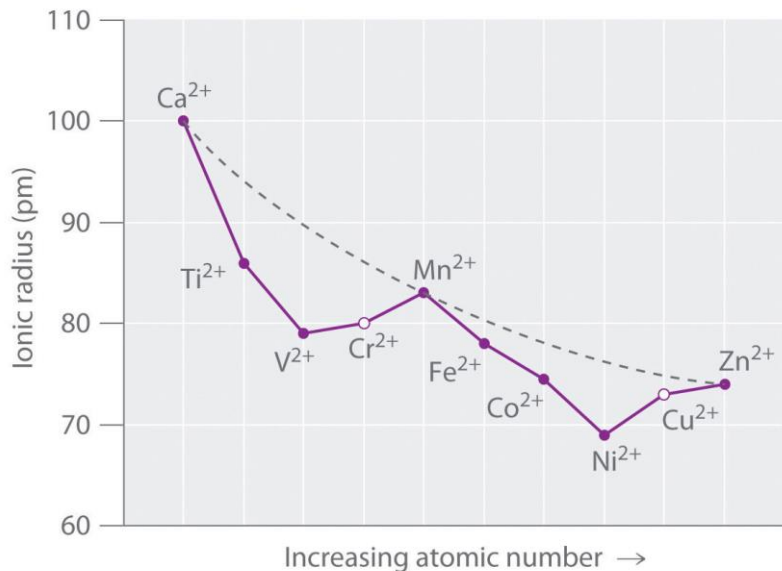
# Periodic Table of the Elements

## Ionic Radius, in picometers

	1																18	
1	H 10																He	
2	Li 90	Be 59										B <sup>3+</sup> 41	C <sup>4+</sup> 30	N 132	O 126	F 119	Ne	
3	Na 116	Mg 86										Al 68	Si <sup>4+</sup> 54	P 212	S 184	Cl 181	Ar	
4	K 138	Ca 100	Sc <sup>3+</sup> 75	Ti <sup>2+</sup> 86	V <sup>2+</sup> 79	Cr <sup>2+</sup> 73	Mn <sup>2+</sup> 83	Fe <sup>2+</sup> 61	Co <sup>2+</sup> 65	Ni <sup>2+</sup> 69	Cu <sup>2+</sup> 73	Zn 74	Ga 62	Ge <sup>4+</sup> 53	As 222	Se 191	Br 195	Kr
5	Rb 149	Sr 118	Y <sup>3+</sup> 90	Zr <sup>4+</sup> 72	Nb <sup>3+</sup> 72	Mo <sup>3+</sup> 65	Tc <sup>5+</sup> 60	Ru <sup>5+</sup> 62	Rh <sup>3+</sup> 67	Pd <sup>2+</sup> 86	Ag 115	Cd 95	In 80	Sn <sup>4+</sup> 69	Sb 245	Te 221	I 220	Xe
6	Cs 167	Ba 135	Lu <sup>3+</sup> 86.1	Hf <sup>4+</sup> 71	Ta <sup>3+</sup> 72	W <sup>6+</sup> 60	Re <sup>5+</sup> 58	Os <sup>6+</sup> 55	Ir <sup>4+</sup> 63	Pt <sup>4+</sup> 63	Au <sup>1+</sup> 137	Hg <sup>2+</sup> 102	Tl 89	Pb <sup>4+</sup> 84	Bi <sup>3+</sup> 96	Po 230	At 227	Rn
7	Fr	Ra	Lr	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn	Nh	Fl	Mc	Lv	Ts	Og
			Lanthanides															
			La <sup>3+</sup> 103	Ce <sup>3+</sup> 101	Pr <sup>3+</sup> 99	Nd <sup>3+</sup> 98	Pm <sup>3+</sup> 97	Sm <sup>3+</sup> 96	Eu <sup>3+</sup> 95	Gd <sup>3+</sup> 94	Tb <sup>3+</sup> 92	Dy <sup>3+</sup> 91	Ho <sup>3+</sup> 90	Er <sup>3+</sup> 89	Tm <sup>3+</sup> 88	Yb <sup>3+</sup> 87		
			Actinides															
			Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No		

Orange colors are anions. Recall that most transition metals are multivalent, so the shown radius is for the oxidation state listed only. Unmarked oxidation number values are the assumed common monovalent oxidation number.

Due to varying oxidation states, graphical representations are usually confined to groups and periods (for transition metals and inner transition metals).



### Ionic Radius Graph

